

Name: _____ Period: _____ Date: _____

Additional Practice Problems

1. Name each of the following compounds:

- a. $\text{Ca}(\text{HSO}_4)_2$ _____
- b. H_3PO_4 _____
- c. HI _____
- d. $\text{Fe}(\text{BrO}_3)_2$ _____
- e. BiAsO_4 _____
- f. $\text{As}_2(\text{SO}_3)_3$ _____
- g. H_2CO_3 _____
- h. $\text{Sn}(\text{NO}_2)_2$ _____
- i. $(\text{NH}_4)_2\text{HPO}_4$ _____
- j. N_2O _____
- k. FeBr_3 _____
- l. HCl _____
- m. HClO _____
- n. NaClO _____
- o. KHCO_3 _____
- p. HNO_3 _____
- q. KMnO_4 _____
- r. HClO_4 _____
- s. Zn_3N_2 _____
- t. PbS_2 _____
- u. PI_5 _____
- v. HNO_2 _____

2. Derive the formula for each of the following compounds:

- a. tin (IV) bromide _____
- b. mercury (II) nitrite _____
- c. sulfurous acid _____
- d. titanium (IV) sulfide _____
- e. iron (III) carbonate _____
- f. iron (II) acetate _____
- g. sodium chromate _____
- h. magnesium oxalate _____
- i. nickel (II) permanganate _____
- j. calcium chlorate _____
- k. bromic acid _____
- l. hydrosulfuric acid _____
- m. manganese (II) hydroxide _____
- n. dinitrogen pentoxide _____
- o. sodium hypochlorite _____
- p. acetic acid _____
- q. chromium (III) sulfite _____
- r. sodium oxalate _____
- s. antimony (III) sulfate _____
- t. potassium cyanide _____
- u. arsenic (V) carbonate _____
- v. lead (II) nitrate _____
- w. hydrofluoric acid _____
- x. ammonium sulfate _____
- y. hypochlorous acid _____