LEARNING REFLECTION SHEET				
CHAPTER 6 OBJECTIVES	**Have you learned this concept or skill (Scale of 1-10)?**	**What Evidence Do You Have that You Learned It? (List at least 2, or more, particular notes, assignments, labs, or quizzes you completed that demonstrate your skill or knowledge)**	**What scores did you receive on the assignments, labs or quizzes you listed?**	**Is this a concept or skill you anticipate needing additional practice with before the test? How do you know?**
SECTION 6-1				
Define chemical bond and explain why most atoms form chemical bonds (i.e. lower potential energy and electrical attraction)				
Describe the basic concepts of ionic and covalent bonding				
Classify bonding type according to electronegativity differences				
VOCAB: chemical bond, ionic bonding, covalent bonding, nonpolar covalent bond, polar, polar covalent bond				
SECTION 6-2				
Define molecule and molecular formula				

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Explain the relationship between potential energy, distance between approaching atoms, bond length, and bond energy.				
State the octet rule				
Be able to write Lewis structures				
Explain how to determine Lewis structures for molecules containing single bonds, multiple bonds, or both				
Explain why scientists use resonance structures to represent some molecules				
VOCAB: molecule, molecular compound, chemical formula, molecular formula, diatomic molecule, bond length, bond energy, octet rule, lone pair, structural formula, single bond, double bond, triple bond, multiple bonds, electron-dot notation, resonance				

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SECTION 6-3				
Compare and contrast chemical formula for a covalent/molecular compound with one for an ionic compound				
Discuss the arrangement of ions in crystals				
Define lattice energy and explain its significance				
List and compare the properties of ionic and covalent/molecular compounds				
Write the Lewis structure for a polyatomic ion				
VOCAB: ionic compound, formula unit, lattice energy, polyatomic ion				

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SECTION 6-4				
Describe the electron-sea model of metallic bonding, and explain why metals are good electrical conductors				
Explain why metal surfaces are shiny and why they are malleable and ductile				
VOCAB: metallic bonding, malleability, ductility				
SECTION 6-5				
Explain VSEPR theory				
Predict the shapes of molecules or polyatomic ions using VSPER theory				
Explain why determines molecular polarity				
VOCAB: VSEPR, names of molecular geometries				