

Name: _____ Date: _____ Period: _____

Section 10-3 - The Gas Laws

Before you leave for the day, be able to identify and explain the difference between a direct and inverse relationship between Pressure, Volume and Temperature of a gas.

Record the amount of each gas species you added: Heavy: _____ Light: _____

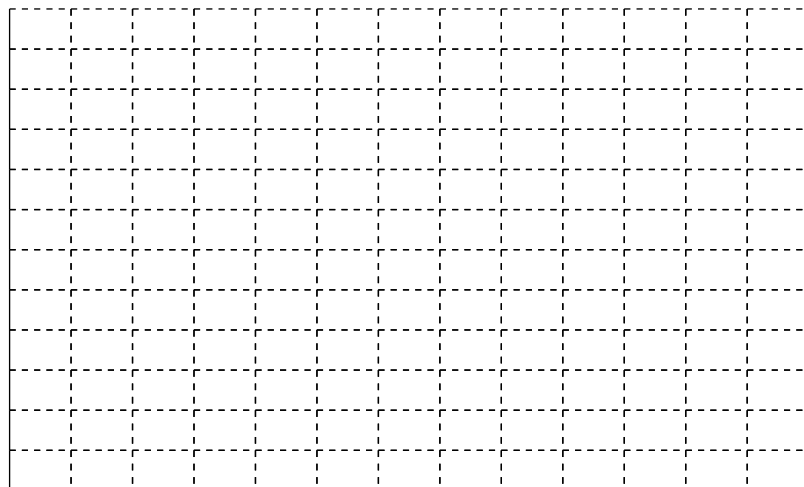
After the gas has been added, hit the pause button and record the pressure: _____

What are the graduations on the ruler? (How much is each small notch worth?) _____

Scenario 1:

	Trial 1	Trial 2	Trial 3	Trial 4	Trial 5
Volume (L)					
Pressure (atm)					

Now with the data you gathered, sketch Pressure(x) vs. Volume(y) in the following graph. Label the graph appropriately. Draw a line best connecting all of the points.



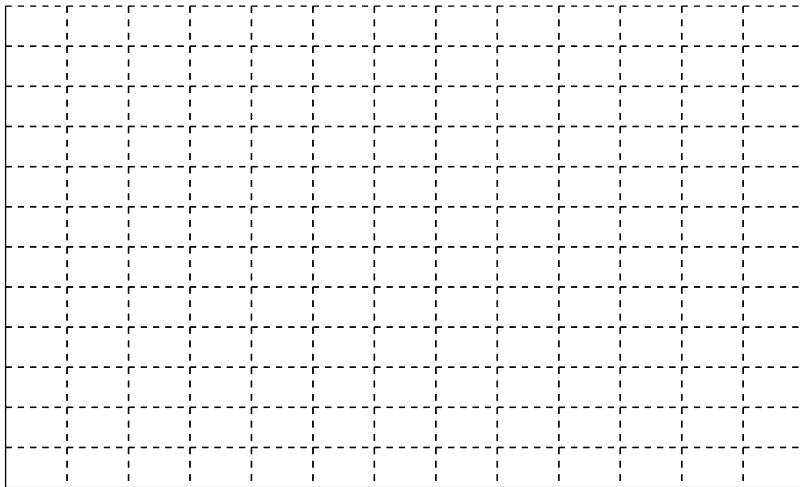
Describe the relationship between Pressure and Volume.
(When one increases, what happens to the other?)

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Scenario 2:

	Trial 1	Trial 2	Trial 3	Trial 4	Trial 5
Volume (L)					
Temperature (K)					

Now with the data you gathered, sketch Temperature (x) vs. Volume (y) in the following graph. Label the graph appropriately. Draw a curved line best connecting all of the points.



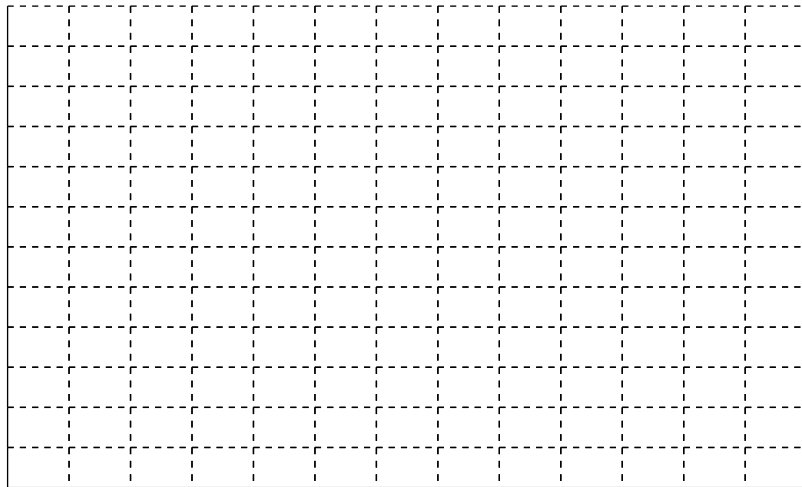
Describe the relationship between Temperature and Volume.
(When one increases, what happens to the other?)

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Scenario 3:

	Trial 1	Trial 2	Trial 3	Trial 4	Trial 5
Temperature (K)					
Pressure (atm)					

With the data you gathered, sketch Temperature (x) vs. Pressure (y) in the following graph. Label the graph appropriately. Draw a line best connecting all of the points.



Describe the relationship between Temperature and Pressure.
(When one increases, what happens to the other?)