

Study Guide for Chapter 7 – see bottom of this page for review assignment

7.1

Explain the significance of a chemical formula.

Determine the formula of an ionic compound formed between two given ions

Name an ionic compound given its formula

Using prefixes, name a binary molecular compound from its formula

Write the formula of a binary molecular compound given its name

7.2

List the rules for assigning oxidation numbers

Give the oxidation number for each element in the formula of a chemical compound

7.3

Calculate the formula mass or molar mass of any given compound

Use molar mass to convert between mass in grams and amount in moles of a chemical compound

Calculate the number of molecules, formula units, or ions in a given molar amount of a chemical compound

Calculate the percentage composition of a given chemical compound

7.4

Compare and contrast empirical and molecular formulas

Determine the empirical formula of a compound using experimentally determined masses

Determine the empirical formula of a compound using percentage composition

Determine the molecular formula of a compound given the empirical formula and the molar mass

What to Study??

- You will not be given the naming flowchart, so study it
 - You need to understand how to identify a compound as ionic, covalent, or an acid and name it appropriately
- KNOW THE POLYATOMIC IONS (p. 210)
 - There will not be specific questions asking what each one is, but you will need to correctly name compounds that contain them
- All past assignments
 - Naming & making chemical formulas
 - Assigning oxidation numbers
 - Molar conversions, percentage composition, molar masses
 - Empirical and molecular formula assignments

REVIEW ASSIGNMENT: #38-42, 44-49 ON PAGE 237-238. DUE TUESDAY 3/11 (DAY OF TEST)

