

Name: _____

Period: _____

Date: _____

Mixed Practice Naming Compounds & Writing Chemical Formulas

(includes ionic compounds, molecular compounds, and acids)

Write the formula for the following compounds.

1. Silicon dioxide _____
2. Carbon tetrachloride _____
3. Tin (IV) chromate _____
4. Barium hydroxide _____

Write the names of the following compounds.

5. PI_3 , stock system _____
6. N_2O_4 _____
7. $\text{Fe}(\text{NO}_2)_2$ _____
8. CCl_4 _____
9. CO _____
10. CuCO_3 _____

Write the formula and give the names of the compounds formed by the following ions.

11. Ca^{2+} and Cl^- _____
12. Pb^{2+} and CrO_4^{2-} _____
13. Al^{3+} and SO_4^{2-} _____
14. Sn^{4+} and PO_4^{3-} _____

Write the formula and indicate the charges for the following ion.

15. Sulfide ion _____

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16. Copper (I) ion _____

17. Carbonate ion _____

Name the following acids.

18. HF _____

19. HBr _____

20. H_2CO_3 _____

Write the formula for the following acids.

21. Phosphoric acid _____

22. Carbonic acid _____

23. Nitrous acid _____

More Practice Naming Compounds

(1 point each)

24. NaBr _____

25. CaSO_4 _____

26. K_2S _____

27. $\text{Ni}(\text{NO}_3)_2$ _____

28. Mg_3N_2 _____

29. $\text{Fe}_2(\text{CO}_3)_3$ _____

30. Cr_2O_3 _____

31. $\text{Ti}(\text{ClO}_4)_4$ _____

32. AlCl_3 _____

33. PbC_2O_4 _____

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34. CO _____

35. CO₂ _____

36. S₂F₆ _____

37. P₄O₁₀ _____

38. N₂O₄ _____

39. NCl₃ _____

40. PBr₅ _____

41. SiS₂ _____

42. N₂F₄ _____

43. SeBr₂ _____

44. Write formulas for each of the following compounds (1 each):

a. barium hydroxide _____

g. disulfur dichloride _____

b. sodium nitrate _____

h. carbon diselenide _____

c. lead (II) sulfite _____

i. acetic acid _____

d. potassium permanganate _____

j. chloric acid _____

e. iron (II) phosphate _____

k. sulfurous acid _____

f. diphosphorous trioxide _____

l. phosphoric acid _____

45. Fill in the blanks in the table with the correct name or chemical formula.

Compound Name	Formula
Aluminum sulfide	1.

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Aluminum sulfite	2.
3.	PbCl_2
4.	$(\text{NH}_4)_3\text{PO}_4$
Hydro-iodic acid	5.

46. Covalently bonded binary compounds typically form from what types of elements (2 points)?

47. Binary ionic compounds typically form from what types of elements (2 points)?

48. What is the best way to tell a binary acid from an oxyacid (2 points)?

49. If you have a metal in your compound, what type of compound do you have and how do you name it (2 points)?