

## Study Guide for Chapter 5 Quiz on Wednesday, January 8, 2014

### Know and be able to apply the following information:

- Know the common names of groups on the periodic table (including alkali metals, alkaline earth metals, transition metals, halogens, and noble gases)
- Understand how the periodic table can be used to predict the electron configuration of an element
- Know what valence electrons are.
- Understand how the periodic table (specifically the groups) help predict how many valence electrons an element will have.
- Understand how and WHY the periodic table (specifically the groups) help predict the common ionic charge an element will have.
- **Understand ionization energy and electronegativity.**
- **Know the trends** across period and down groups for ionization energy and electronegativity.
- Given an element, be able to identify its period, block, and common charge using a periodic table.
- Write electron configurations for outer shell (valence) electrons

### Study:

- Packet of notes for section 5.3 for trends of the periodic table.
- Notes from the board about valence electrons and common charges